

ERRATA FOR 2ND PRINTING OF WATER CHEMISTRY

1. (2/22/05) Inside back cover: the final value shown for $RT_{298.15K}$ should be 24.79 L-bar/mol (not 24.17 L-bar/mol).
2. (2/25/04) p.3. In Table 1.1., the concentration of Mg^{2+} in seawater should be given as 1280 mg/L, not 128 mg/L.
3. (1/16/02) p.30. In Table 1.4a, in the explanation of the terms in the Extended Debye-Huckel equation, the expression for B should be: $B = 50.3(\epsilon T)^{-1/2}$ (in the book, the minus sign is missing from the exponent).
4. (10/01/03) p.77, second line from bottom: delete the word “and” after the comma.
5. (10/25/03) p.79, Equation 2.30: the left hand side should be $\frac{dPE_{tot}}{dx}$.
6. (10/25/03) p.110, first line: the reference to Equation (2.79) should be to Equation (2.80).
7. (10/08/03) p.111, Example 2.10. In the fourth line of the solution to part a (the second equation shown), the left hand side of the equation should be $\Delta\bar{G}_r^{\circ}$ instead of $\Delta\bar{G}_r$.
8. (9/23/03) p.111-113, Example 2.10. In the problem statement, in part b , lines 1-2, replace the phrase “a system,” with “1.0 L of a solution.” At the end of that sentence, add the sentence: Assume that the molar volumes of H_2O , $HCrO_4^-$, and $Cr_2O_7^{2-}$ are all equal (1 L per 55.56 mol), so that the solution volume remains constant as the reaction proceeds.

On the first line of p.112, replace “Assuming that” with “Because.”

Revise footnote 13 as follows: Although the activity of water is very nearly 1.0 regardless of the distribution of Cr species, the number of moles of H_2O in the solution changes slightly when some of the Cr converts from one form to the other. As a result, the contribution of H_2O to G_{tot} changes as the reaction proceeds, and this change must be taken into account.

Revise Equation (2.92) by placing the V on the right hand side:

$$G_{\text{tot}} = V \left[c_{\text{H}_2\text{O}} \left(\bar{G}_{\text{H}_2\text{O}}^{\circ} + RT \ln \frac{c_{\text{H}_2\text{O}}}{\sum c_i} \right) + c_{\text{HCrO}_4^-} \left(\bar{G}_{\text{HCrO}_4^-}^{\circ} + RT \ln c_{\text{HCrO}_4^-} \right) \right. \\ \left. + c_{\text{Cr}_2\text{O}_7^{2-}} \left(\bar{G}_{\text{Cr}_2\text{O}_7^{2-}}^{\circ} + RT \ln c_{\text{Cr}_2\text{O}_7^{2-}} \right) \right] \quad (2.92)$$

In the second line below Equation (2.92) replace “water” with “total species.”

Revise Equation (2.94) as follows:

$$a_{\text{H}_2\text{O}} = \frac{55.56 - c_{\text{HCrO}_4^-} - c_{\text{Cr}_2\text{O}_7^{2-}}}{55.56} = \frac{55.56 - c_{\text{HCrO}_4^-} - 0.5(0.10 - c_{\text{HCrO}_4^-})}{55.56} \\ = \frac{55.51 - 0.5c_{\text{HCrO}_4^-}}{55.56} \quad (2.94)$$

In the second line below Equation (2.94), delete “per unit volume of solution.” In the third and fourth lines below the equation, change “ G_{tot}/V ” to “ G_{tot} .”

Replace Figure 2.8 and its caption with the following figure (note: the values on the y axis have changed, and the y axis label now represents G_{tot} , not G_{tot}/V).

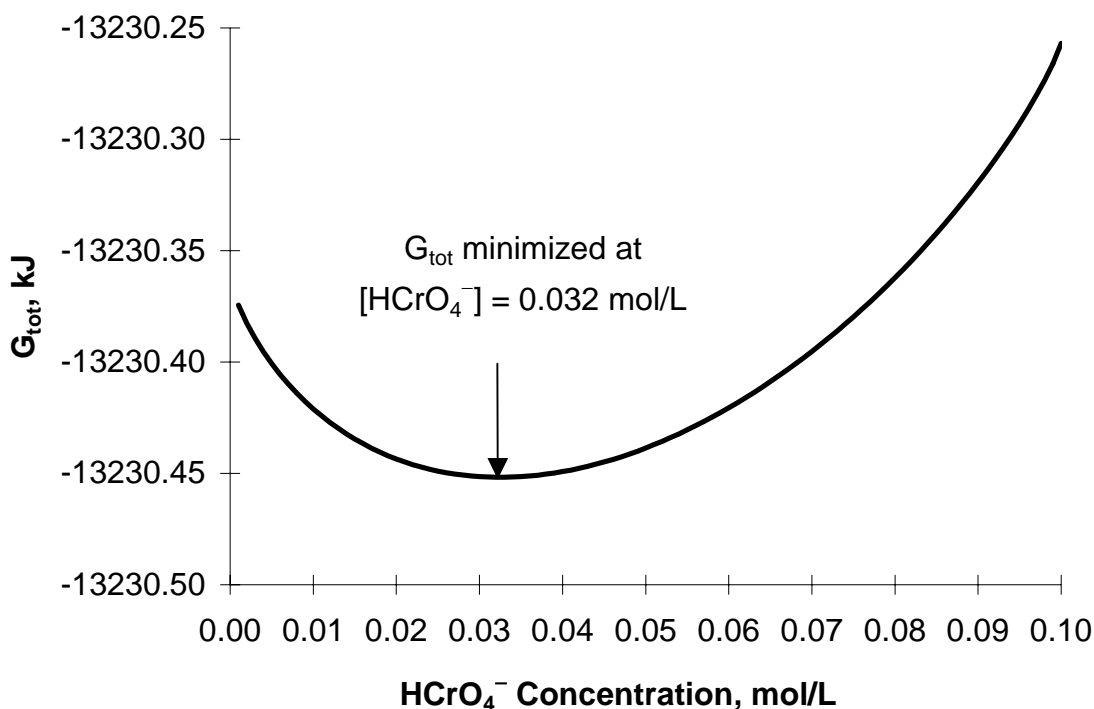


Figure 2.8 Total Gibbs energy of 1.0 L of a solution containing $10^{-1} M$ $c_{\text{Cr,tot}}$, as a function of the HCrO_4^- concentration.

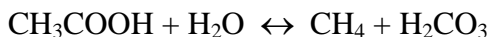
9. (10/25/03) p.116, first equation in solution to Example 2.12: replace 16.40 with 16.32.
10. (10/25/03) p.124, solution to Example 2.14: In last equation in part *b* and both equations in part *c*, replace 3.03 with 2.96, and 1.74 with 1.72.
11. (10/11/05) p.128, Problem 3. At the end of the paragraph preceding part *a*, add: Reaction 1 is at equilibrium, but Reaction 2 might not be.
12. (11/13/03) p.135: In Equation 3.5, change 10^{-1} to 1.
13. (10/25/03) p.136, second line of caption to Figure 3.3: replace Al^{3+} with Cu^{2+} .
14. (11/13/03) p.139, Table 3.2: Add the value 13.41 for $\text{p}K_{a3}$ for arsenous acid (H_3AsO_3). Make the same change in the copy of this table on p.642.
15. (2/14/03) p.160, line 2. “undergoes *k*” should be “undergoes *n*”
16. (10/25/03) p.168, first line of last paragraph: replace “nine” with “seven.”
17. (11/12/03) p.184-185. On the bottom line of the table on p.184, the pH of the solution with $[\text{NaCl}] = 0.1 M$ should be 3.30. Also, add a line to the bottom of that

table, for the species “log [H⁺]”. On this line, below the values of 3.26 and 3.30 on the “pH” line, put values of –3.26 and –3.19, respectively.

Revise paragraph at bottom of p.184 and top of p.185 as follows: The water with NaCl has slightly higher H⁺ and F[–] concentrations than the salt-free solution, because the salt decreases the activity coefficients of these two species. Since the activity coefficient of HF is 1.0 in both solutions, the decreases in γ_{H^+} and γ_{F^-} force more HF to dissociate to satisfy the K_a expression, and this extra dissociation releases more H⁺ and F[–] to solution. Correspondingly, the concentration of HF that remains undissociated declines. Despite the increase in the concentrations of H⁺ and F[–], the activity of each of these species declines, because their activity coefficients decline. The resulting, fractional decline in the product {H⁺} {F[–]} is exactly equal to the fractional decline in {HF}, so the ratio {H⁺} {F[–]}/ {HF} is the same as in the solution with no NaCl, and equal to K_a . Because {H⁺} is lower in the salty solution, pH is higher, as indicated. The effect of the salt would be greater if more highly charged species were involved (for example, CO₃^{2–} or PO₄^{3–}).

18. **(9/25/03)** p.187. In question 5, the correct formula for ascorbic acid is C₅H₉O₄–COOH.
19. **(10/16/03)** p.195. In the equation just above the horizontal line, the fraction and final result should be $\frac{(10^{-7.95})(10^{-3.00})}{10^{-6.05}} = 10^{-4.90}$.
20. **(10/26/05)** p199, part a, last equation: should be “(where {H⁺} = {OCl[–]})” instead of (where {H⁺} + {OCl[–]})”.
21. **(10/25/03)** p.202, final line: Last two terms in the equality should be 10^{1.77} and 58.9, instead of 10^{–1.77} and 0.017.
22. **10/21/03)** p.217-218. In the table near the bottom and the equation immediately below it, the values 1.0x10^{–10} and 1.0x10^{–4} should be replaced by 7.94x10^{–11} and 1.26 x 10^{–4}, respectively. Also, in the table, the values 0.56 x 10^{–4} and 0.44 x 10^{–4} should be reversed.
- In addition, on the last line of the page, the value –1.56x10^{–4} should be replaced by –1.82x10^{–4}. The same replacement should be made on p.218 in the third-to-last line of the example, and the value 1.12x10^{–4} in the final two lines should be replaced by 1.13x10^{–4}.
23. **(10/28/05)** p.229. In the table, in the row with the Cu(NH₃)_x²⁺ species, the entry under –2 should be Cu(NH₃)₄²⁺, and the entry under +1 should be Cu(NH₃)₂²⁺.
24. **(1/27/03)** p.231. Most of question 4a is irrelevant, since the values were inadvertently printed on the axes.

25. (10/25/03) p.249, solution to Example 5.2: In the fourth line of the second paragraph, replace 2.67×10^{-3} and 0.67 with 3.32×10^{-3} and 0.83, respectively. Later in the paragraph, replace 4.92×10^{-3} with 5.57×10^{-3} . Make the corresponding changes (3.32 for 2.67 and 5.57 for 4.92 in the summation at the end of the example.
26. (10/25/03) p.254, fourth line from bottom: replace “large” with “wide.”
27. (10/28/03) p.255. In the last line on the page, H_2CO_2 should be H_2CO_3 .
28. (10/28/05) p.264. In the table, the entries for Cl^- and HCl should not be italicized.
29. (10/25/03) p.272. In the second line of the *Solution*, part *a*, replace “base” with “acid.”
30. (10/27/03) p.284. In part *b* of the answer to Example 5.10, the reference shown to Table 5.5 should actually be to Table 5.6.
31. (10/28/03) p.284. In part *b*, in the equation for β_{PO_4} , the value 10^{-3} should be 10^{-2} .
32. (10/28/03) p.285. In the table at the top of the page, β_{NH_3} at pH 9.25 (the last entry in the last column) should be 1.73×10^{-2} .
33. (10/28/03) p.286. In part *b* of the answer to Example 5.11, the reference shown to Table 5.5 should actually be to Table 5.6.
34. (10/28/05) p.288. Problem 1(a)(vi): $\text{p}K_{a1}$, $\text{p}K_{a2}$, and $\text{p}K_{a3}$ should be 3.13, 4.72, and 6.33, respectively.
35. (2/13/03) p.290. In problem 9, the first reaction shown should be:



Part *b* of the question should be reworded as follows: What would the solution pH be if all the sulfite were destroyed by the second reaction, and then all the acetic acid that remained was destroyed by the first reaction?

36. (2/13/03) p.290. In problem 11, line 4, the word “strong” should be inserted before the word “acid.”
37. (11/08/04) p.292. In Problem 14, second paragraph, second sentence: replace “contains negligible Ca^{2+} and alkalinity” with “can be represented as a mixture of water, H_2NOM , and NaCl”.
38. (10/31/03) p.299: In the chemical reaction shown at the bottom of the page, the coefficient 1 for H^+ should be underlined.
39. (10/31/03) p.301: In the table near the top of the page, the heading **Stoichiometric Coefficient** should be centered over the three columns labeled

- H₂O**, **OH⁻**, and **HAc**, and the underline under that heading should extend only over those three columns.
40. **(11/13/03)** p.332. In the footnote to Table 7.2, change 20°C to 25°C. Make the same change in the copy of the table on p.643.
 41. **(2/7/03)** p.334. In the fourth line of the caption for Figure 7.2, change “that concentration of Cl⁻ and” to “the concentration of Cl⁻ indicated, and”
 42. **(11/13/03)** p.336: In the first equation in the solution to Example 7.3, in both places where °C appears in the denominator of the fraction, it should be change to K (Kelvins).
 43. **(10/21/02)** p.341. In the equation just above 7.22, the coefficient “2” preceding {HAc} should be deleted.
 44. **(2/7/03)** p.342. The second paragraph in the Solution to Example 7.6 should begin, “Rather than write a charge balance...”
 45. **(11/10/03)** p.347. In the PC table, the first two values in the right hand column should be 2.96×10^{-5} and 4.44×10^{-5} .
 46. **(11/10/03)** p.348. In Equation (7.23), the subscript on the first {OH⁻} in the equation should be changed from fii to fin.
 47. **(11/13/03)** p.354: In the table at the bottom, the equilibrium concentration of HCO₃⁻ should be 1.95×10^{-3} .
 48. **(11/13/03)** p.358: In problem 8c, the mass balance should state that the rate at which sulfur enters the column (moles of S per unit time), considering both inlet streams, equals the rate at which it leaves, considering both outlet streams, i.e.,

$$(Q_{\text{gas,in}} c_{\text{S,gas,in}} + Q_{\text{aq,in}} c_{\text{S,aq,in}} = Q_{\text{gas,out}} c_{\text{S,gas,out}} + Q_{\text{aq,out}} c_{\text{S,aq,out}} .$$
 49. **(11/13/03)** p.361: In Problem 16, line 3: change Example to Problem.
 50. **(11/21/03)** p.368: In Table 8.2, Log β₃ for Hg²⁺ should be 32.16, and Log *β₃ should be -9.84. The same corrections are needed on the copy of Table 8.2 on p.644.
 51. **(11/22/04)** p.375. In Table 8.3, the values in the NH₃ column in the rows corresponding to Zn²⁺ should be 2.2, 4.5, 7.0, and 9.0 for ZnL, ZnL₂, ZnL₃, and ZnL₄, respectively. The same changes should be made in the copy of Table 8.3 on p.646.
 52. **(12/05/04)** p.398. In Table 8.7, add the following entries to the table: SiO₂(s), log K_{s0} = -2.74; ZnO(s), log K_{s0} = -16.12. Also, Log K_{s0} for Hematite (α-Fe₂O₃) should be -81.26, and Log K_{s0} for Hydroxyl apatite (Ca₅(OH)(PO₄)₃(s)) should be -58.2. The same changes should be made to the copy of Table 8.7 on p.647.
 53. **(12/04/03)** p.418, 3rd line of final paragraph: change 10^{-13.73} to 10^{-14.73}.
 54. **(11/29/03)** p.420, line 9: change Figure 8.10 to Figure 8.12.

55. (2/27/03) p.454. In problem 4, line 3, Cu should be Cu^{2+} . Also, in part *c*, add the words, “showing all species that are present at $>10^{-6} M$ ” to the end of the question.
56. (2/25/03) p.455. In Problem 7, for Species B, the coefficient for NH_4^+ should be 0, and the log K should be 6.35.
57. (12/05/04) p.456. In Problem 11, right before the last sentence, add: Ignore hydrolysis of Ca^{2+} .
58. (9/23/03) p.456. In Problem 13b, before the parenthetical remark about hydrolysis reactions, add the statement: Ignore formation of Al-SO₄ complexes.
59. (3/5/03) p.459. In the second paragraph of Problem 19, replace the wording at the end of the problem, starting with the sentence that begins “Carry out the calculations...” as follows: Carry out the calculations twice, using first one and then the other of the following two assumptions. (1) The rain is in equilibrium with the atmosphere before striking the monument, but no more gases dissolve as the rain drips down the monument’s surface; (2) The rainwater is in continuous equilibrium with the atmosphere as the water drips.

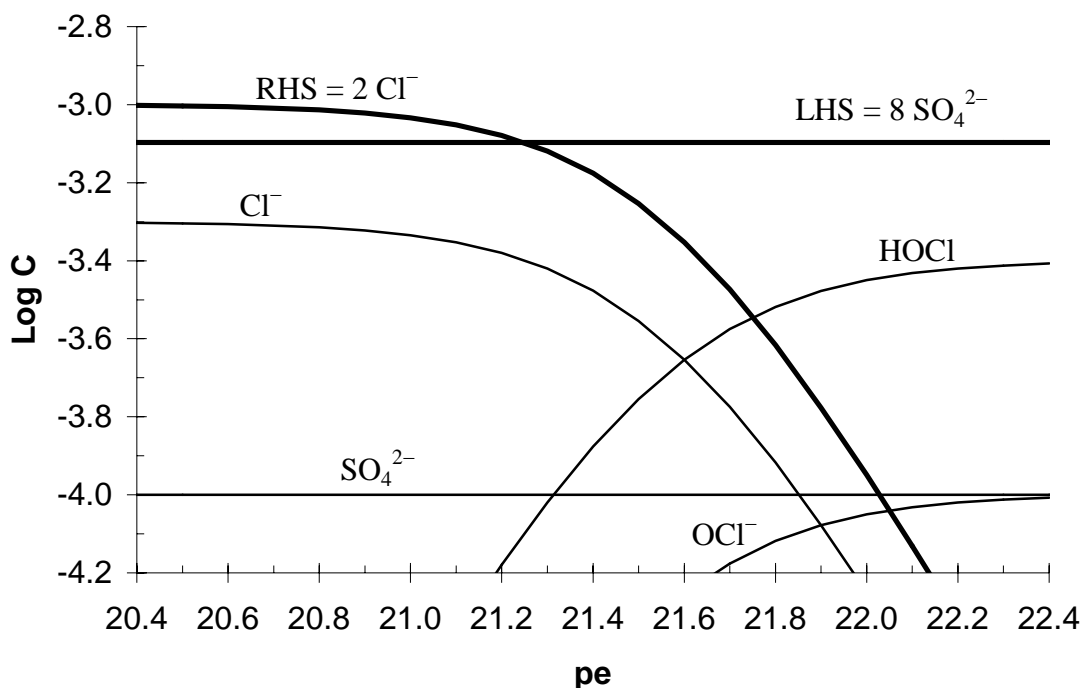
Also, at the end of the problem, add the following. Hint: for the systems that do not equilibrate continuously with the atmosphere, make a guess for the amount of solid that dissolves, and iterate on that guess until it is consistent with the charge balance.

60. (12/1/04) p.459. In Problem 20c, at the end of the first sentence, add:
 $K_{s0} = 10^{-2.74}$.
61. (12/2/03) p.473: in the solution to Example 9.4, part *a*, in the last line of text, delete the word ‘is’.
62. (12/10/03) p.487 and 499. At the bottom of Figures 9.4 and 9.5, the line for the $\text{H}^+/\text{H}_2(\text{g})$ system should say pH 10.5, not pH 7.
63. (12/2/03) p.500. In Table 9.5, entry #5, in the expression for Δpe , the final term should not have RT in the numerator; the expression should be as follows (note: the same correction is needed in the copy of Table 9.5 on page 652):

$$\Delta\text{pe} = \Delta\text{pe}^\circ - \frac{1}{n_e} \log Q$$

64. (12/2/03) p.501, lines 2-3: The term \bar{G}_e^- should be on one line.
65. (3/10/03) p.505. In the last equation on the page, above the final two paragraphs of text, the left hand side of the equation should be $8\text{pe}^\circ(\text{W})$, not 8pe° .

66. (3/10/03) p.508. Figure 9.9 is missing curves for HOCl and OCl⁻. The correct figure is shown below.



67. (3/5/03) p.546. At the end of Problem 2, before the parenthetical remark, add the sentence: Consider HOCl, OCl⁻, and Cl⁻ as the only significant Cl species.
68. (3/5/03) p.546. In Problem 3, paragraph 2, *TOTHg* should be *TOTHg(II)*.
69. (3/10/03) p.547. At the end of part a in Problem 6, add the sentence: Assume that the acid/ base reactions in the system reach equilibrium quickly.
70. (12/19/01) p.549. Add the following statement to problem 11, before the parenthetical note at the end: Ignore redox reactions of water.
71. (4/28/05) p.596. The fractions in Equations 10.47 and 10.48 should all be inverted. The same fractions appear in the bottom of Table 10.6 on p.604, and the same corrections should be made to them.
72. (10/28/02) p.604. The minus sign should be removed from the exponential argument in the equation for $K_{A,int,o}$.
73. (4/15/05) p.607. The third and fourth equations in the middle of the page are both missing a term equal to $\{H^+\}^{-1}$ on the right-hand side. This term should appear in the middle of the equation, after the term $\{A^{2+}\}$.

74. **(12/05/04)** p.647. Add values of -25.55 for $\log K_{s0}$ of $\text{HgO}(s)$ and -39.28 for $\log K_{s0}$ of $\text{Hg}(\text{CN})_2(s)$ to Table 8.7. Also, $\text{Log } K_{s0}$ for Hematite ($\alpha\text{-Fe}_2\text{O}_3$) should be -81.26 , and $\text{Log } K_{s0}$ for Hydroxyl apatite ($\text{Ca}_5(\text{OH})(\text{PO}_4)_3(s)$) should be -58.2 .